6 Reactions in Aqueous Solutions

Ionic Compound

Ions in Solution

Na^+		Cl⁻		NaCl	NaCl
				Ag^+	NO
			Na^+	NO	NaNO ₃
Ag^+	Cl⁻	AgCl	$NaNO_3$		

Solubility Rules for Some Common Salts in Water

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double displacement reactions

AB + XY AY + XB

А

X B

Y

6.2 Acids and Bases

• H⁺ • OH⁻

HCl, HBr, HI, HNO $_3$, H $_2$ SO $_4$, and HClO $_4$

6.2.1 Dissociation of Acids

____ H⁺

 H_2SO4 (aq) H^+ (aq) + HSO

6.3 Oxidation Reduction Reactions

 $\begin{array}{cccc} 2 \text{ Na } (s) \ + \ Cl_2 \ (g) & 2 \text{ NaCl } (s) \\ \text{Na} & Cl & \\ & \text{Na}^+ & Cl^- \end{array}$

Na Na⁺ + e^{-} Cl + e^{-} Cl⁻

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Oxidation

Reduction

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6.4.2 Synthesis Reactions

6.4.3 Decomposition Reactions

decomposition reactions